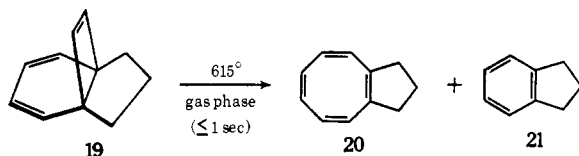


octene intermediate (e.g., **9**) will depend upon the capability of the trigonal cyclobutene carbons to bend in the direction of the neighboring diene unit (or *vice versa*). Molecular models indicate that a trimethylene bridge substantially restricts this requisite motion. Not unexpectedly, therefore, the pyrolysis of **19**¹⁸ required enhanced temperatures and gave only low yields of **20** and **21** in which the fragmentation product **21** predominated by a factor of 2.



This work also raises the intriguing possibility that cyclooctatetraene and certain of its derivatives may be capable of skeletal rearrangement in their own right. We are currently exploring this question.

Acknowledgment. Appreciation is expressed to the National Cancer Institute for partial support of this research. The sample of **19** was kindly supplied by Mr. Ronald K. Russell whom we thank.

(18) L. A. Paquette, R. E. Wingard, Jr., and R. K. Russell, *J. Amer. Chem. Soc.*, in press.

(19) National Institutes of Health Predoctoral Fellow, 1969–1971.

Leo A. Paquette,* Robert E. Wingard, Jr.¹⁹

Department of Chemistry, The Ohio State University
Columbus, Ohio 43210

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Additions and Corrections

A Study of the Decay Processes in the Triplet Sulfur Dioxide Molecule Excited at 3828.8 Å [*J. Amer. Chem. Soc.*, **93**, 2587 (1971)]. By HOWARD W. SIDEBOTTOM, CHARLES C. BADCOCK, JACK G. CALVERT,* GEORGE W. REINHARDT, BLAINE R. RABE, and EDWARD K. DAMON, Department of Chemistry and the ElectroScience Laboratory, The Ohio State University, Columbus, Ohio 43210.

On page 2588, column 2, line 15, the length of the tube should read 88 cm. Also, on page 2592, column 1, line 8 and in the caption to Figure 6, the pathlength should read 88 cm.

On page 2592, the sentence beginning on line 20 should read: The slopes give: $\epsilon_{\text{SO}_2} = 0.11 \pm 0.02$ and $\epsilon_{\text{Ac}_2} = 8.3 \pm 0.8 \text{ l./mol cm}$ at 3828.8 Å.

An Equilibrium Centrifugation Study of the Self-Association of *N*-Methylacetamide in Carbon Tetrachloride Solutions at 25° [*J. Amer. Chem. Soc.*, **93**, 7075 (1971)]. By ROBERT J. ALBERS, ANNE B. SWANSON, and GORDON C. KRESHECK,* Department of Chemistry, Northern Illinois University, DeKalb, Illinois 60115.

The values given in line 5 of the abstract should read 73 g/mol, $2.68 \pm 0.15 \times 10^4 \text{ ml/g}$, $0.00312 \pm 0.00036 \text{ mol-ml/g}^2$, and $0.00970 \pm 0.00201 \text{ mol-ml}^2/\text{g}^3$, respectively.

In the column headings of Table I, the units of BM_1 should read $(\text{g/ml})^{-1}$ and those of CM_1 should read $(\text{g/ml})^{-2}$.

Photolysis of Ortho-Substituted Aryl Azides in Diethylamine. Formation and Autoxidation of 2-Diethylamino-1*H*-azepine Intermediates [*J. Amer. Chem. Soc.*, **94**, 513 (1972)]. By RICHARD J. SUNDBERG,* STUART R. SUTER, and MARTIN BRENNER, Department of Chemistry, University of Virginia, Charlottesville, Virginia 22901.

On page 518, column 1, the final sentence is in error. The structure under discussion, 2,7-dimethyl-4*H*-

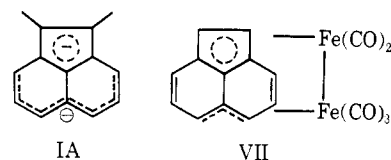
azepin-4-one, was later retracted: E. Bullock, B. Gregory, and A. W. Johnson, *J. Chem. Soc.*, 1632 (1964).

Structure and Reactivity of Organic Ions in Gas-Phase Radiolysis. V. The Structure and Formation Process of the C_7H_7^+ Ions from Toluene [*J. Amer. Chem. Soc.*, **94**, 661 (1972)]. By YUKIO YAMAMOTO,* SETSUO TAKAMUKU, and HIROSHI SAKURAI, The Institute of Scientific and Industrial Research, Osaka University, Suita, Osaka, Japan.

In footnote 13, the feature of the nmr spectrum is incorrect and should be H_{Ph} , τ 2.9~3.1; H_{CH_3} , τ 6.1; and H_{CH_2} , τ 7.8.

Deprotonation of Acenaphthene. The Acenaphthylene Dianion [*J. Amer. Chem. Soc.*, **94**, 985 (1972)]. By L. D. KERSHNER, J. M. GAIDIS, and H. H. FREEDMAN,* The Dow Chemical Company, Eastern Research Laboratory, Wayland, Massachusetts 01778.

Structures Ia and VII should be



Photoisomerization about the Carbon-Nitrogen Double Bond of an Oxime Ether [*J. Amer. Chem. Soc.*, **94**, 1000 (1972)]. By ALBERT PADWA* and FRED ALBRECHT, Department of Chemistry, State University of New York at Buffalo, Buffalo, New York 14214.

Due to a mathematical error in the determination of the light intensity, the quantum yield reported in the last two lines of page 1000 should be $\Phi = 0.30$ and $\Phi = 0.36$ at 0.036 *M*.

On the Mechanism of Intermolecular Aromatic Substitution by Arylnitrenes [*J. Amer. Chem. Soc.*, **94**,